

Read Reaction Ratios



Use the matching sheet ID when checking answers.

Use the material to work through each task. Show the requested response.

1. What is the $\text{H}_2:\text{O}_2$ mole ratio?

Answer: 2:1

2. How many moles of O_2 are needed for 6 moles of H_2 ?

Answer: 3 moles

3. With enough oxygen, how many moles of H_2O form from 4 moles of H_2 ?

Answer: 4 moles

4. There are 4 moles of H_2 and 1 mole of O_2 . Which reactant limits the modeled reaction?

Answer: O_2

5. How many moles of H_2O can that mixture form?

Answer: 2 moles

6. Why is changing a coefficient different from changing a subscript?

Look for:

- Keeps chemical identity distinct from quantity.

One possible response: A coefficient changes amount; a subscript changes the substance's composition.

Open responses use criteria and an illustrative response. Other well-supported responses can be acceptable.